



## ACROSS

- 1 Electron \_\_\_\_\_ is the process by which an electron moves from one atom or molecule to another atom or molecule.
- 7 A \_\_\_\_\_ is a compound containing an oxygen-oxygen single bond.
- 8 A terminal electron \_\_\_\_\_ is a compound that receives or accepts an electron during cellular respiration or photosynthesis.
- 9 A \_\_\_\_\_ reaction is either the oxidation or reduction reaction component of a redox reaction.
- 10 The oxidation \_\_\_\_\_ is an indicator of the degree of oxidation of an atom in a chemical compound. It is usually numerically equal to the oxidation number.
- 11 A \_\_\_\_\_ agent is the element or a compound that reduces another species in a redox reaction. It is the electron donor in the redox.
- 12 In general terms, an electron \_\_\_\_\_ gives up an electron during cellular respiration.

## DOWN

- 2 \_\_\_\_\_ is the free radical product of the one-electron reduction of dioxygen.
- 3 A \_\_\_\_\_ agent is a substance that gains electrons in a redox chemical reaction.
- 4 \_\_\_\_\_ is a bimolecular process involving the exchange of bonds between the two reacting chemical species, which results in the creation of products with similar or identical bonding affiliations.
- 5 Within a molecule or complex, the \_\_\_\_\_ number of an element is the charge that it would have if the compound were composed of ions, with assignment of shared electrons based on electronegativity.
- 6 Standard reduction \_\_\_\_\_ is the tendency of a chemical species to acquire electrons and thereby be reduced.
- 11 \_\_\_\_\_, shorthand for reduction/oxidation reaction, is a term used to describe chemical reactions in which atoms have their oxidation state changed.